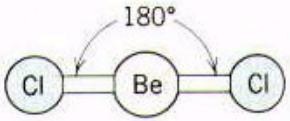
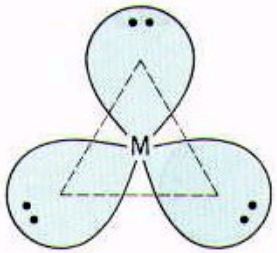
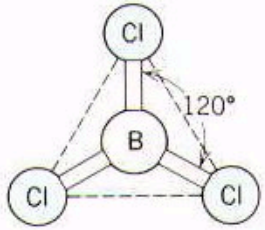
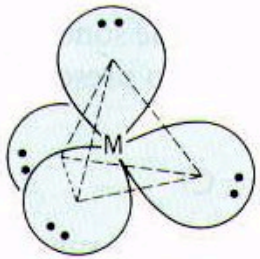
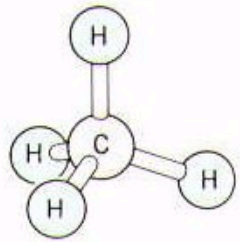
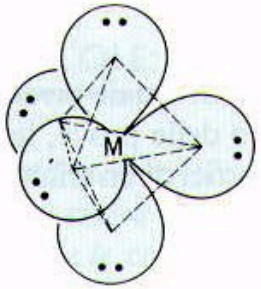
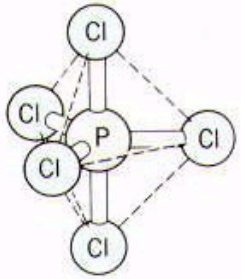
Numero di coppie di elettroni	Forma		Esempio
Due coppie		Lineare	$BeCl_2$ 
Tre coppie		Triangolare planare	BCl_3 
Quattro coppie		Tetraedrico (Un tetraedro ha quattro facce triangolari e quattro angoli).	CH_4 

Cinque coppie

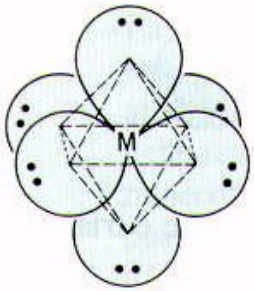


Trigonale bipiramidale
(Questa figura consiste di due piramidi unite per condivisione di una faccia. Il piano triangolare attraversa il centro).

PCl_5

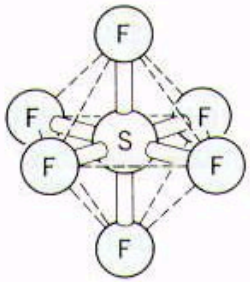


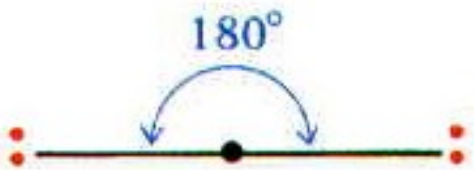
Sei coppie



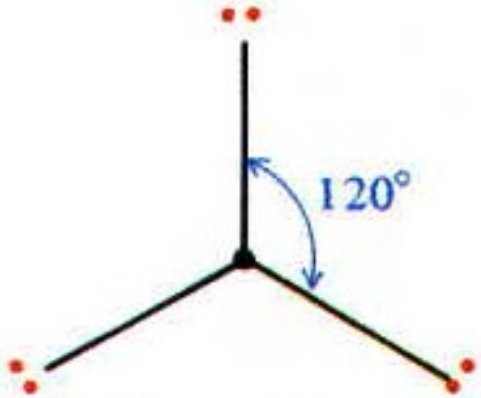
Ottaedrica
(Un ottaedro è una figura a otto facce e *sei* angoli. Esso è formato da due piramidi quadrangolari che hanno la base in comune).

SF_6

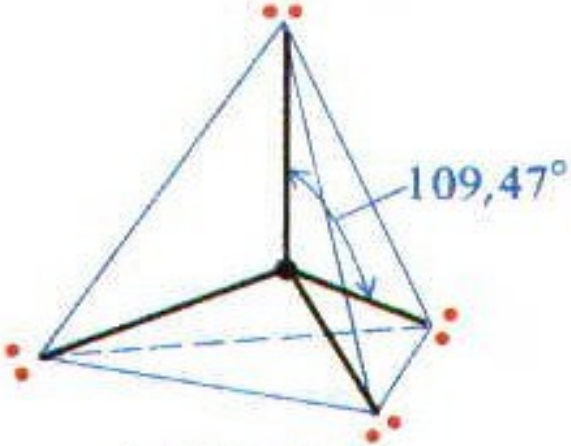




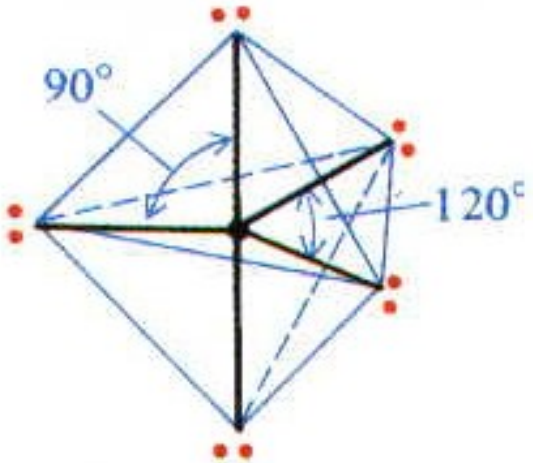
2: Lineare



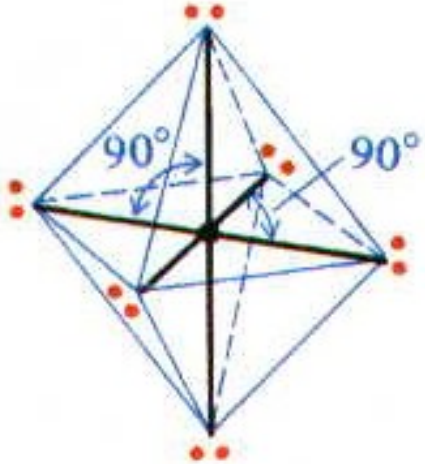
3: Planare trigonale



4: Tetraedrica



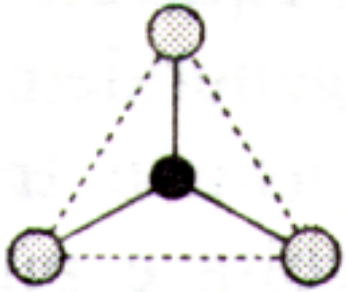
5: Bipiramidale trigonale



6: Ottaedrica

Lineare e trigonale piana

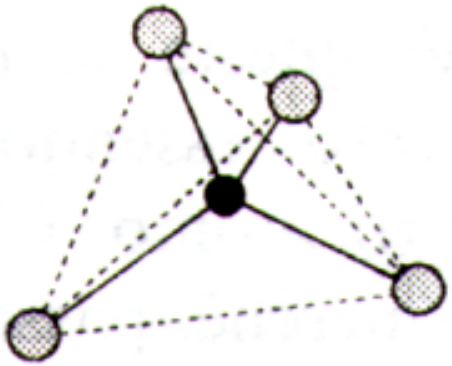
XY_2
lineare



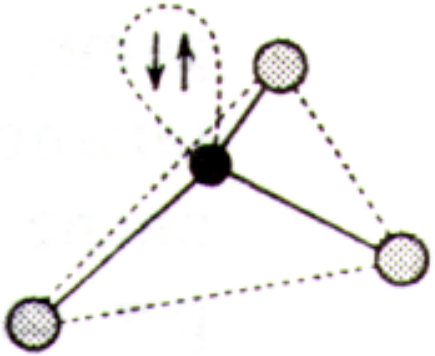
a: XY_3
trigonale piana

b: XY_2
angolare

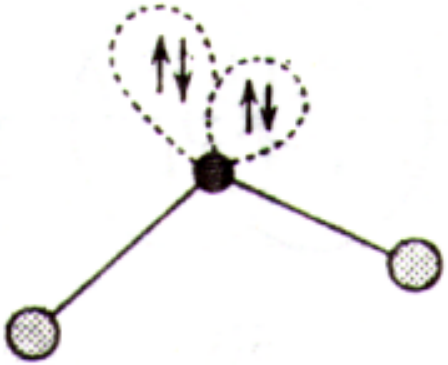
Tetraedrica



a: XY_4
tetraedrica

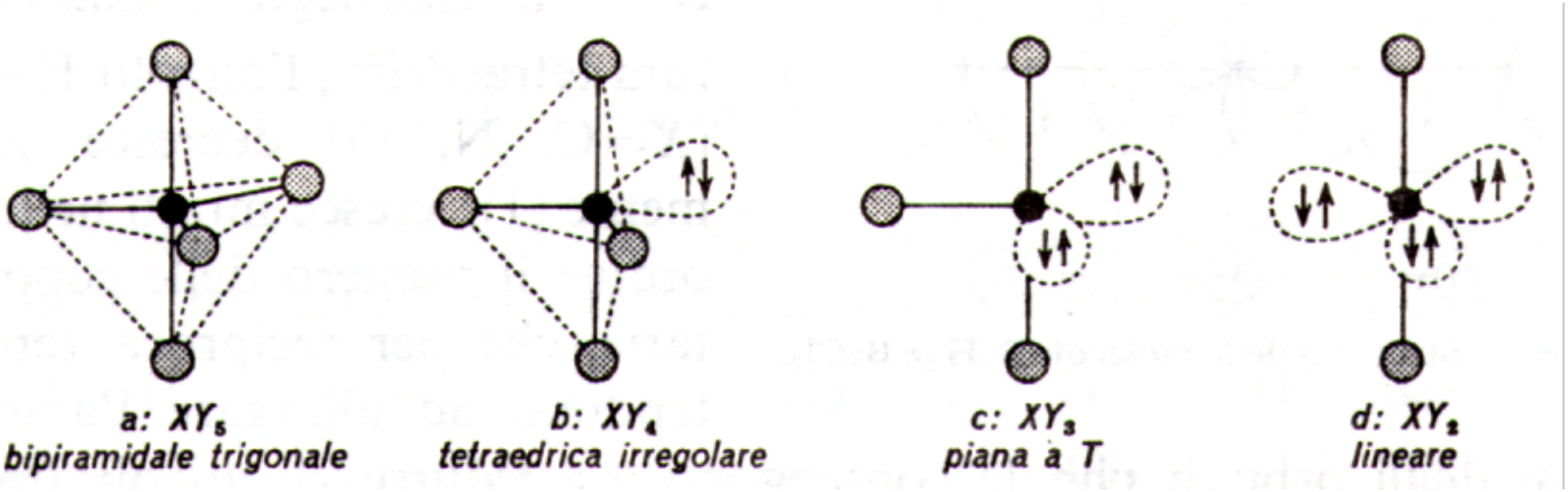


b: XY_3
piramidale trigonale

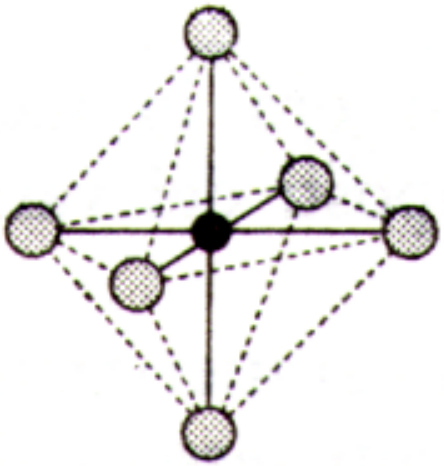


c: XY_2
angolare

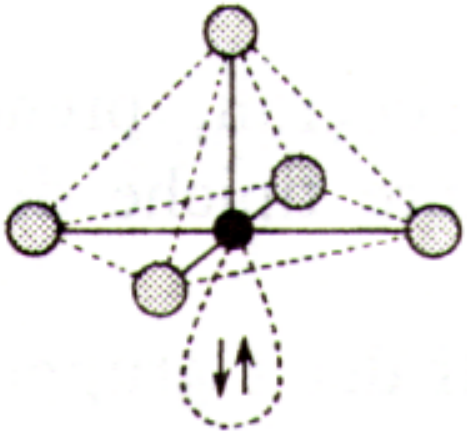
Bipiramide trigonale



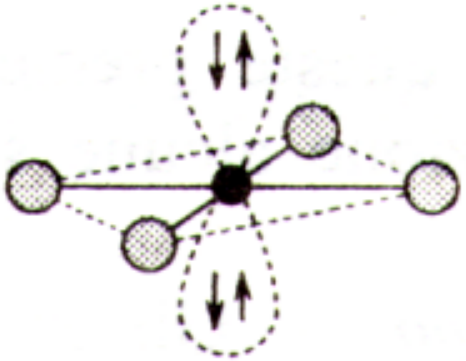
Ottaedrica



a: XY_6
ottaedrica

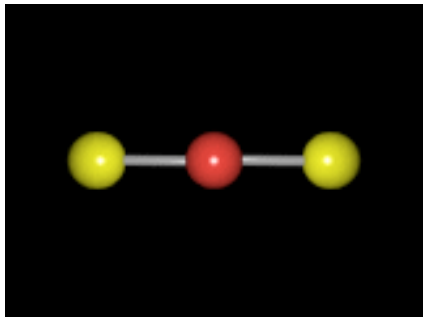


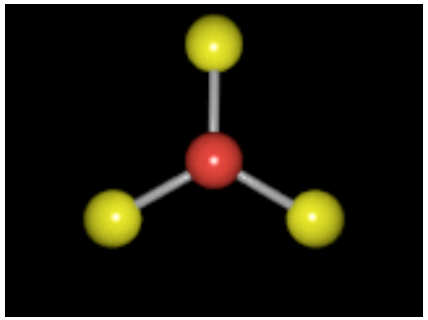
b: XY_5
piramidale tetragonale

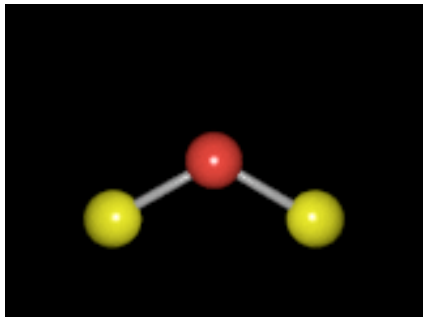


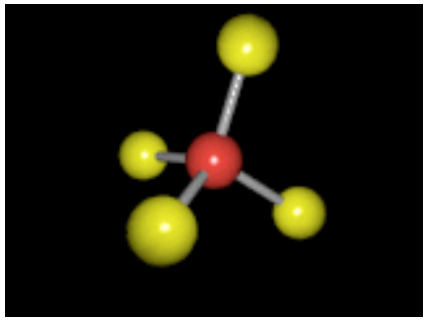
c: XY_4
quadrata piana

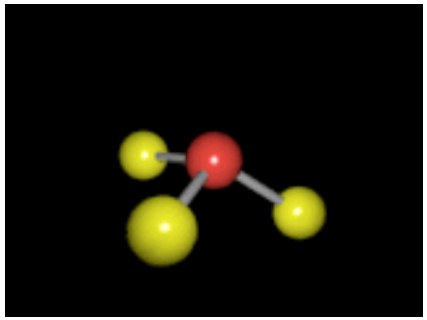
Geometrie

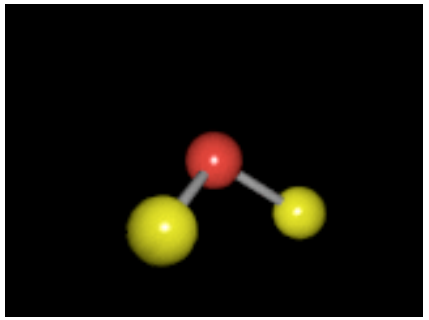




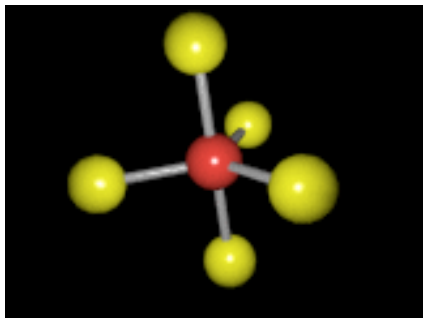
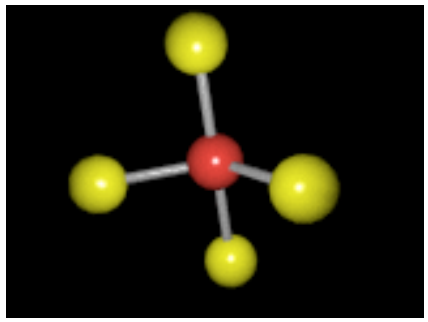
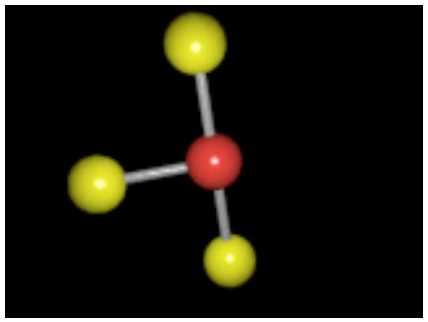
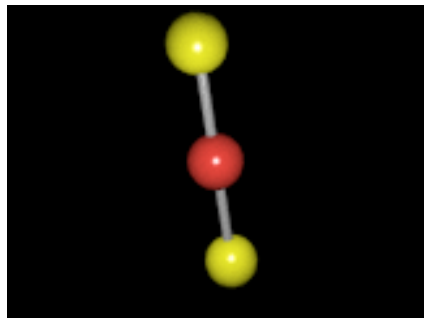
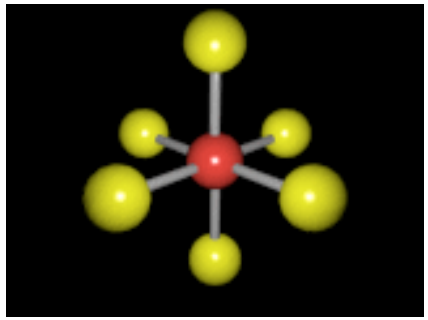
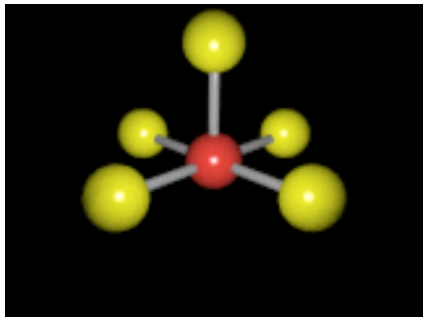
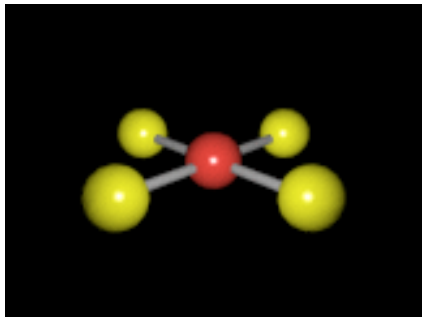


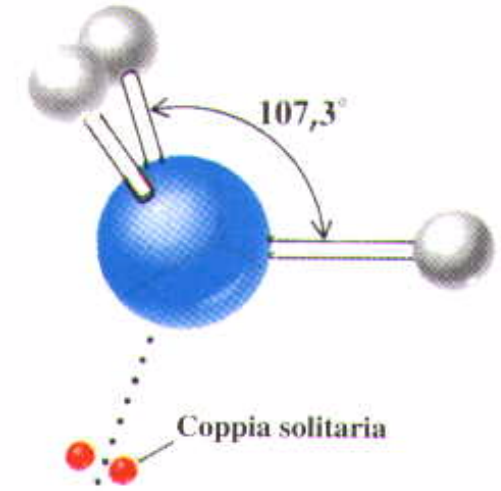


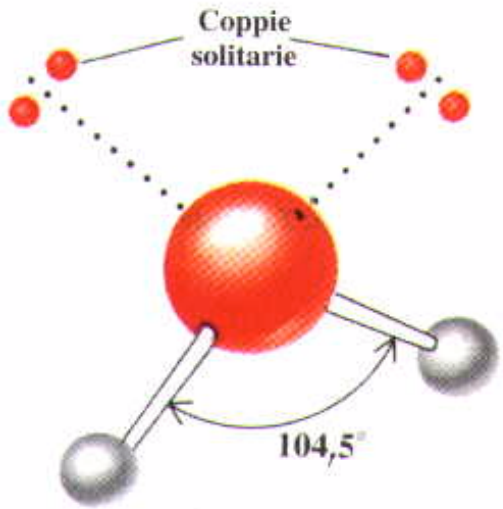


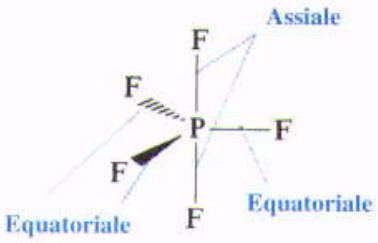


Geometrie

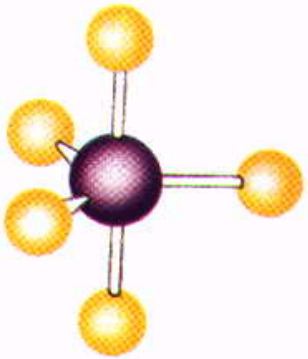









107,3°
Coppia solitaria

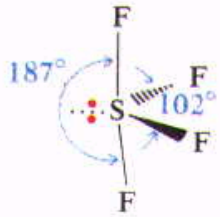

Coppie
solitarie
104,5°



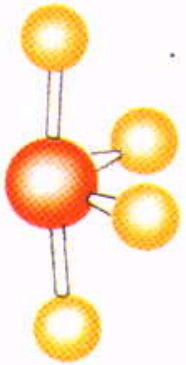
(*a*) PF_5



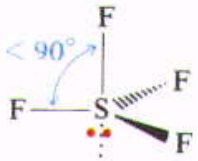
Bipiramide trigonale



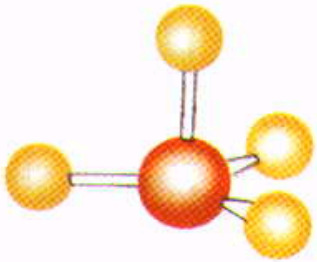
(*b*) SF_4



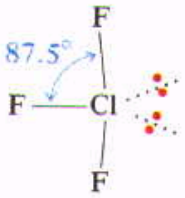
Altalena
(bassa energia, favorita)



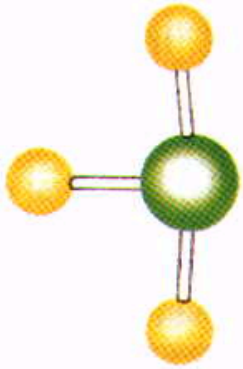
(*c*) SF_4



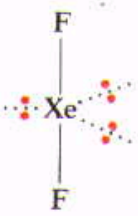
Piramide distorta
(elevata energia, sfavorita)



(*d*) ClF_3



T distorta



(*e*) XeF_2

Lineare